

MUSTAQIL ISHLASH UCHUN MASALALAR YECHIMLARI

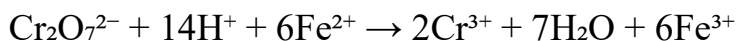
1-masala yechimi

Cr: $+6 \rightarrow +3$ (har bir Cr $3e^-$ qabul qiladi, jami $6e^-$)

Yarim reaksiyalar:

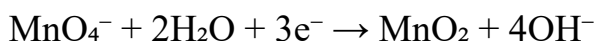
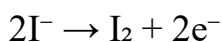


Fe reaksiyasini 6 ga ko'paytiramiz va qo'shamiz:

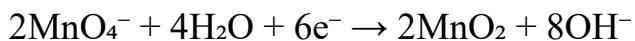


2-masala yechimi

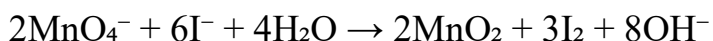
Yarim reaksiyalar (ishqoriy muhit):



Elektronlarni tenglashtiramiz ($6e^-$):



Yig'amiz:



3-masala yechimi

Cu: $0 \rightarrow +2 \rightarrow 2e^-$ beradi $\rightarrow$ qaytaruvchi

N: $+5 \rightarrow +4 \rightarrow 1e^-$ qabul qiladi $\rightarrow$ oksidlovchi (NO_3^-)

4-masala yechimi

S^{2-} da S: -2

SO_4^{2-} da S: $+6$

O'zgarish: $-2 \rightarrow +6$

Elektronlar soni: $8e^-$ beriladi

5-masala yechimi

Mn: $+7 \rightarrow +2 \rightarrow z = 5$

$$E = M / z = 158 / 5 = 31.6 \text{ g}$$

6-masala yechimi

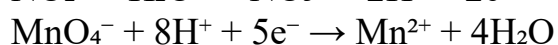
$$n(\text{Al}) = 5.4 / 27 = 0.20 \text{ mol}$$



$$n(\text{e}^-) = 0.20 \cdot 3 = 0.60 \text{ mol e}^-$$

7-masala yechimi

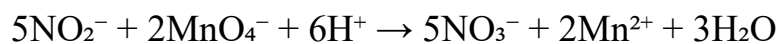
Yarim reaksiyalar:



Elektronlarni tenglashtiramiz (10e^-):



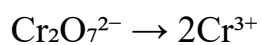
Yig'amiz va qisqartiramiz:



8-masala yechimi

MnO_4^- (kislotali muhit) eng kuchli oksidlovchi,
chunki Mn +7 holatdan +2 gacha qaytariladi va 5e^- qabul qiladi.

9-masala yechimi



Jami elektronlar: 6e^- (1 mol uchun)

0.10 mol uchun:

$$n(\text{e}^-) = 0.10 \cdot 6 = 0.60 \text{ mol e}^-$$

10-masala yechimi

H_2S da S: -2

SO_4^{2-} da S: +6

O'zgarish: -2 $\rightarrow$ +6

Elektronlar soni: 8e^- beriladi

